

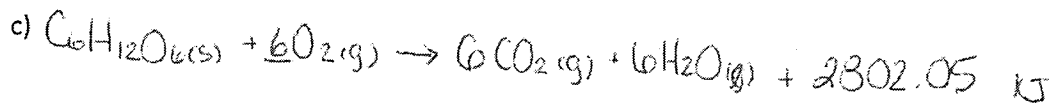
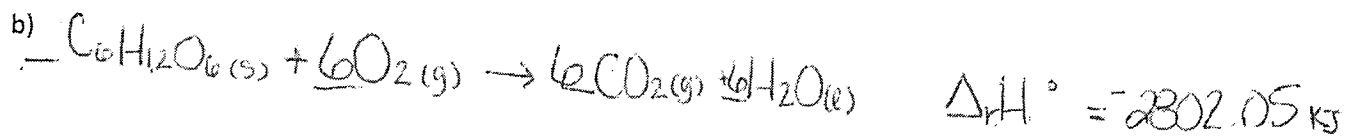
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### Communicating Enthalpy Changes

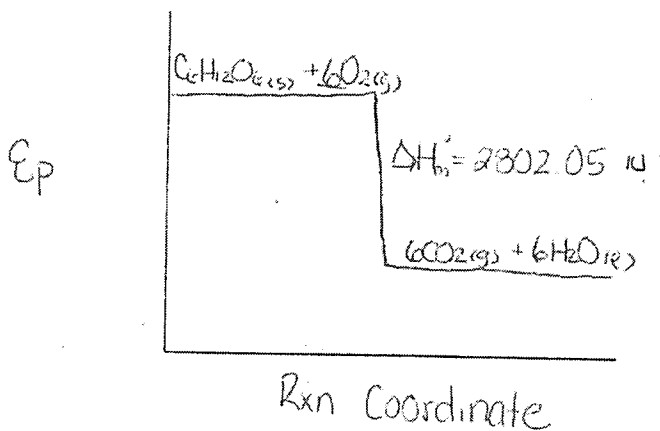
Express the standard enthalpy changes for the following reactions by using the four different methods of communication. Start with a balanced equation.

1. One mole of glucose is consumed, during cellular respiration, to release 2802.05 kJ of energy.

a)  $\Delta_r H_m^\circ = -2802.05 \text{ kJ/mol}$   
 $\text{C}_6\text{H}_{12}\text{O}_6$



d)



2. Ammonium chloride is mixed with water. The energy required for the reaction is 114 kJ/mol.

a)  $\Delta_r H_m^\circ = +114 \text{ kJ/mol}$   
 $\text{NH}_4\text{Cl}$

