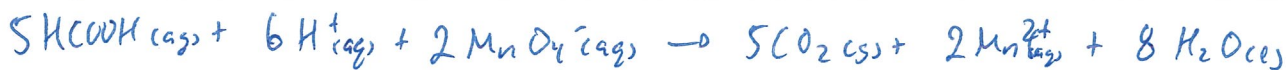
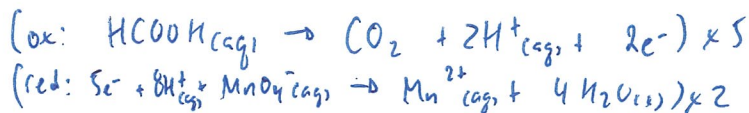
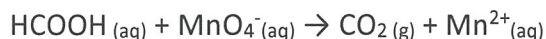
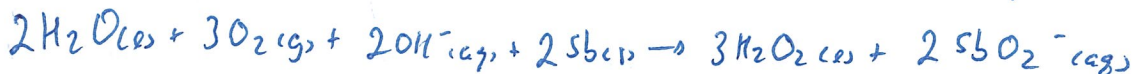
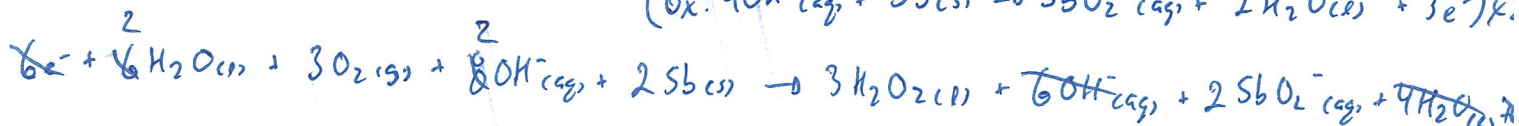
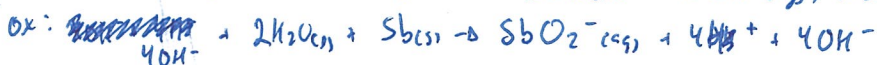
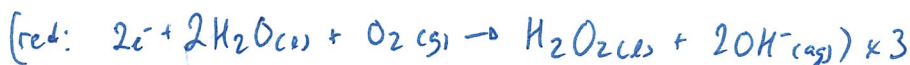
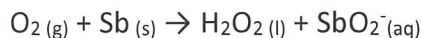


MORE REDOX PRACTICE

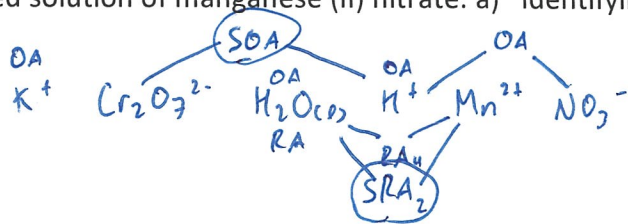
1. Balance the following redox reaction equation, which takes place in an acidic solution, using the half-reaction method.



2. Balance the following redox reaction equation, which takes place in a basic solution, using the half-reaction method.



3. Potassium dichromate solution is mixed with an acidified solution of manganese (II) nitrate. a) identifying all of the species
 b) label them as agents
 c) predict the most likely reaction to occur
 d) write out the respective half-reactions
 e) write the net redox reaction
 f) state whether it will be spontaneous or non-spontaneous.



SPONT.