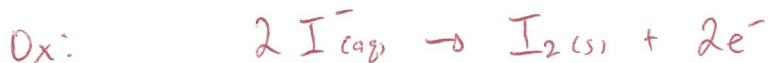
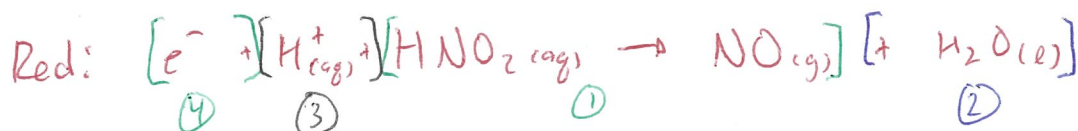


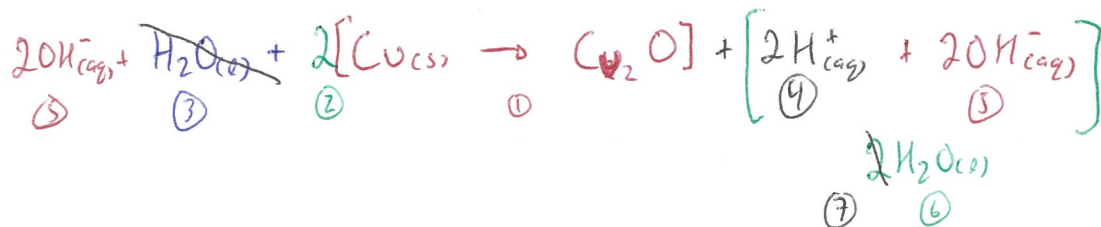
3. Write and label two balanced half-reaction equations to describe the reaction of chlorine gas with aqueous sodium iodide. Identify the OA and RA in the net REDOX equation.



4. Create and label the half-reaction for the following: nitrous acid in an acidic solution will form nitrogen monoxide gas.



5. Create and label the half-reaction for the following: copper metal in a basic solution will form copper (I) oxide.



6. Aqueous iron (III) chloride can react with aqueous sodium hydroxide to produce solid iron (III) hydroxide and aqueous sodium chloride. Has a redox reaction taken place. Explain your answer.



No Redox.  $\text{Fe}^{3+} \rightarrow \text{Fe}^{3+}$   
 $\text{Na}^+ \rightarrow \text{Na}^+$  No transfer of  $\text{e}^-$ , no redox  
 $\text{Cl}^- \rightarrow \text{Cl}^-$   
 $\text{OH}^- \rightarrow \text{OH}^-$